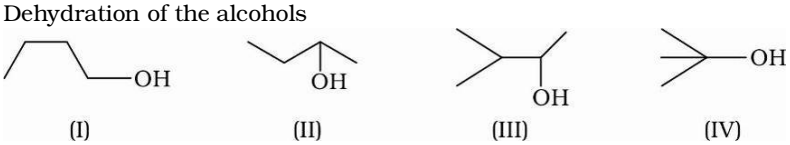
Date Planned : __ / __ / __	Daily Tutorial Sheet-7	Expected Duration : 30 Min
Actual Date of Attempt : __ / __ / __	Level-2	Exact Duration :_____

86. Dehydration of the alcohols



(I) (II) (III) (IV)

will be in order :

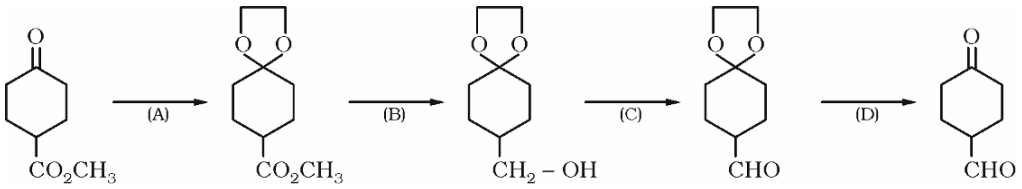
(A) IV > II > III > I **(B)** I > II > III > IV **(C)** IV > II > III > I **(D)** II > IV > I > III

87. When *o*- or *p*-Phenolsulphonic acid is treated with bromine water, the product formed is :

(A) 2, 4-Dibromophenol **(B)** 2, 4, 6-Tribromophenol

(C) 3-Bromophenol **(D)** 3, 5-Dibromophenol

***88.**

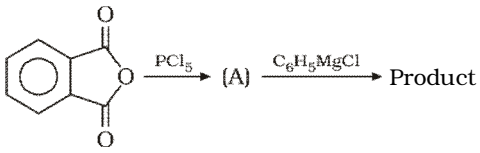


Identify correct combination :

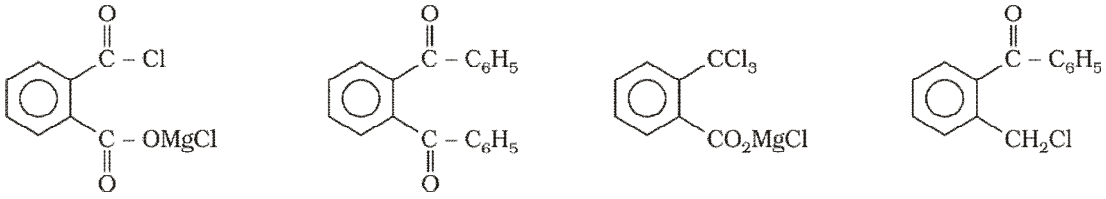
(A) (A) = $\begin{matrix} CH_2-OH \\ | \\ CH_2-SH \end{matrix}$ **(B)** (B) = $LiAlH_4$ / Ether

(C) (C) = $KMnO_4$ **(D)** (D) = $H_3O^{\oplus}$

89.



The product formed in the above reaction is :



(A) **(B)** **(C)** **(D)**

90. $R-\overset{O}{\overset{\|}{C}}-O-H+R_1-OH \overset{H^+}{\rightleftharpoons} R-\overset{O}{\overset{\|}{C}}-O-R_1$

In above esterification reaction, the rate of reaction is maximum, when R_1-OH is :

(A) 1° alcohol **(B)** 3° alcohol

(C) 2° alcohol **(D)** CH_3OH

91. A and B in the given reaction are :

$$B \xleftarrow[H^+,\,\Delta]{KMnO_4} (CH_3)_3COH \xrightarrow{KMnO_4 / OH^-} A$$

(A) A and B both are $(CH_3)_2C = CH_2$

(B) A and B both are $(CH_3)_2CO + CH_2O$

(C) A is $(CH_3)_3COH$ while B is $(CH_3)_2C = CH_2$ or $(CH_3)_2CO$

(D) A and B both are $(CH_3)_3COH$, i.e., there is no reaction

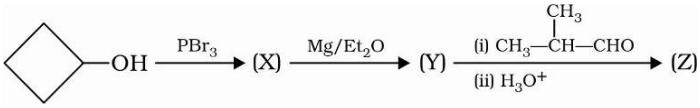
***92.** Which of the following statement(s) is(are) true ?

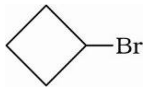
(A) All alcohols are weaker acids than water except methanol

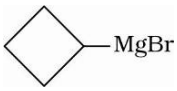
(B) All alcohols are stronger acid than water

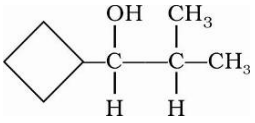
(C) All alcohols are much weaker acids than terminal alkynes

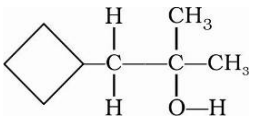
(D) All alcohols are stronger acids than terminal alkynes

***93.** 

(A) X is 

(B) Y is 

(C) Z is 

(D) Z is 

***94.** Lucas test is used to make distinction between 1°, 2° and 3° alcohols

$$ROH + \underset{\text{conc.}}{HCl} \xrightarrow[\text{anhydrous}]{ZnCl_2} \underset{\text{White turbidity}}{RCl\downarrow} + H_2O$$

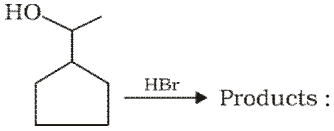
This shows that

(A) ROH behaves as a base.

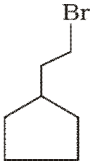
(B) greater the value of pK_a (alcohol), greater the reactivity with conc. HCl and thus sooner the formation of white turbidity.

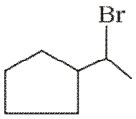
(C) alcohol which reacts fastest with Na metal, will give turbidity at fastest rate.

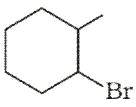
(D) alcohol which gives red colour during Victor Meyer test, will always give turbidity at slower rate than those giving blue or white colour during Victor Meyer test.

***95.** 

Which of the following are possible products in the above reaction ?

(A) 

(B) 

(C) 

(D)